

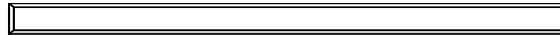


## Exercise 2 (Regenerative Rankine Cycle)

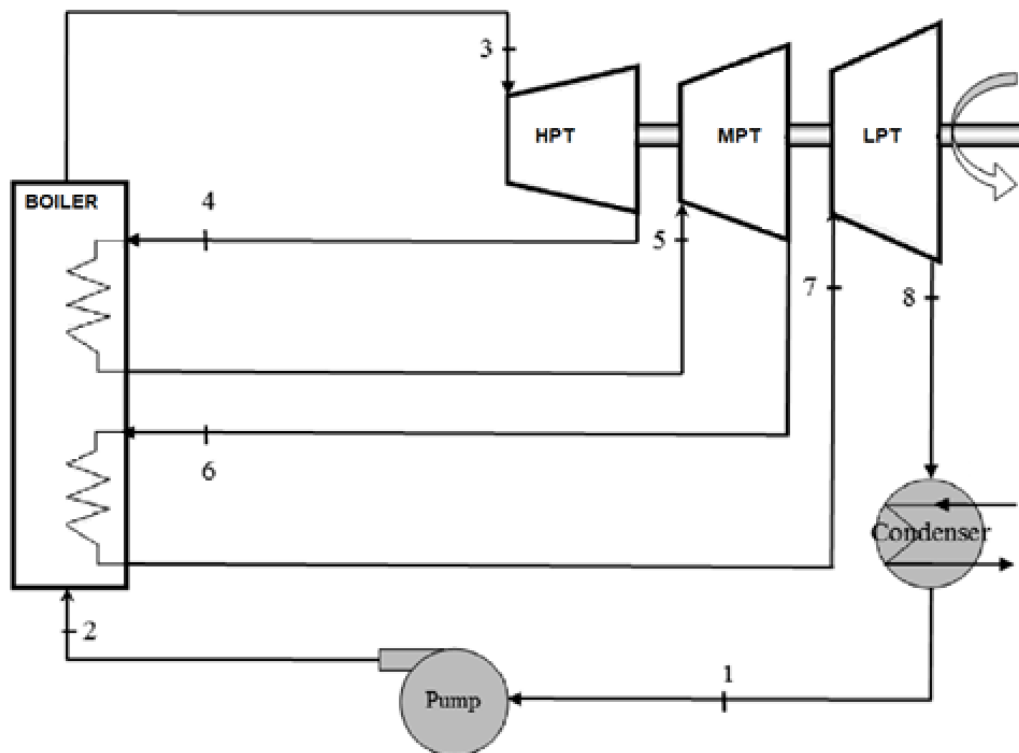
A regenerative Rankine cycle consists of three expansion stages (17000 kPa, 10000 kPa, 2000 kPa, 11 kPa). The gaseous fuel consists of a volumetric mixture of methane (92%) and ethane (8%). Air and fuel enter at 100 kPa and 25°C. The excess air is 160%. The flue-gas leaves the chimney at 427°C and 100 kPa. The relative humidity of the air is 70%.

Get:

- a) Mass balance
- b) Energy balance
- c) Exergy balance



The installation scheme is:

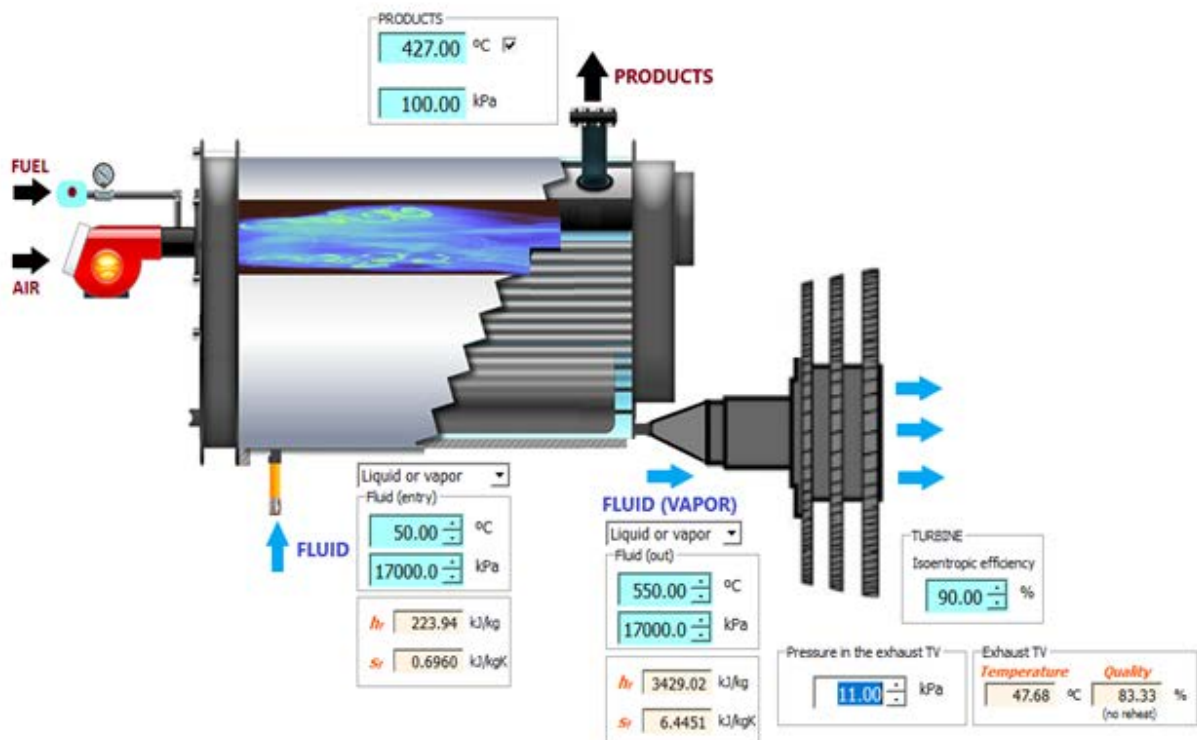




<b>FUEL</b> 25 °C 100.0 kPa kg/s <input type="checkbox"/> m³/s <input type="checkbox"/>		<b>Air Excess (<math>\lambda</math>)</b> 1.600000 <input type="button" value="i"/> 160.000 % <input checked="" type="checkbox"/>	
<b>DRY AIR (feeds CC)</b> 25 °C 100.0 kPa kg/s <input type="checkbox"/> m³/s <input type="checkbox"/>		<b>Air Relative Humidity</b> <input type="button" value="i"/> 70.000 % <input checked="" type="checkbox"/> <b>WB Temperature</b> 20.938 °C <input type="checkbox"/>	

<b>Regenerative Cycle</b>			
<b>Double reheat</b>		Pressure	10000.00 kPa
<input type="button" value="i"/>	(MP turbine inlet)	Temperature	550.00 °C
		Pressure	2000.00 kPa
<input type="button" value="i"/>	(LP turbine inlet)	Temperature	550.00 °C
<input type="button" value="About"/>			

### External Combustion Turbine





**Inerts:**  %

**FUEL**

	$P_{\text{vapor}}(T^* \text{ fuel})$ kPa	Molar ratio %	Chemical Formula	Common Name
Gas T > Tc	<input type="text"/>	<input type="text" value="92.0000"/>	CH <sub>4</sub>	Methane
Gas T < Tc	<input type="text" value="4190.33"/>	<input type="text" value="8.0000"/>	C <sub>2</sub> H <sub>6</sub>	Ethane

Total: 100.0000

	Lower Heating, LHV	Higher Heating, HHV
<b>N</b>	<input type="text" value="852.35"/> kJ/mol fuel	<input type="text" value="943.89"/> kJ/mol fuel
<b>E</b>	<input type="text" value="49656.61"/> kJ/kg fuel	<input type="text" value="54989.64"/> kJ/kg fuel
<b>T</b>	<input type="text"/>	<input type="text"/>

**GROSS**

### Fuel Properties

<b>Enthalpy of formation</b>	<input type="text" value="-75.636"/> kJ/mol fuel	<input type="button" value="i"/>
<b>Absolute entropy (1 atm &amp; 25°C)</b>	<input type="text" value="189.6352"/> J/mol fuel K	<input type="button" value="i"/>
<b>Standard chemical exergy</b>	<input type="text" value="886.44"/> kJ/mol fuel	<input type="button" value="i"/>
<b>Standard Gibbs Free Energy of Formation</b>	<input type="text" value="-49.36"/> kJ/mol fuel	<input type="button" value="i"/>

☒ Experimental correlation  
Kotas (1995)

☒ Thermocombustion database

### Moisture in Air

**Air**

<input type="text" value="0.02269"/>	mol water / mol dry air
<input type="text" value="0.00041"/>	kg water / mol dry air
<input type="text" value="0.78657"/>	mol water / kg dry air
<b>Absolute humidity</b>	
<input type="text" value="0.01417"/>	kg water / kg dry air
<b>Vapor pressure</b>	<b>Dew Point T<sup>a</sup></b>
<input type="text" value="2.21895"/> kPa	<input type="text" value="18.87"/> °C
<b>Air/fuel ratio</b>	
<input type="text" value="16.51893"/>	mol/mol fuel
<input type="text" value="27.53358"/>	kg/kg fuel

### Moisture in Air

**Fuel**

<input type="text" value="0.36655"/>	mol water / mol fuel
<input type="text" value="0.00660"/>	kg water / mol fuel
<input type="text" value="21.35462"/>	mol water / kg fuel
<b>Absolute humidity</b>	
<input type="text" value="0.38471"/>	kg water / kg fuel
<b>Vapor pressure</b>	<b>Dew Point T<sup>a</sup></b>
<input type="text" value="2.21895"/> kPa	<input type="text" value="18.87"/> °C
<b>Air/fuel ratio</b>	
<input type="text" value="16.51893"/>	mol/mol fuel
<input type="text" value="27.53358"/>	kg/kg fuel



Air/Fuel ratio (dry air) AIR

<i>Theoretical</i>	<i>Actual</i>	
10.09524	16.15238	<input type="checkbox"/> mol/mol fuel
16.96804	27.14887	<input type="checkbox"/> kg/kg fuel

Reactants (fuel + dry air)

17.15238 mol/mol fuel

(Assume 1 mol fuel is being combusted)

**FLUE GAS (Combustion Products)**

☒ On wet basis ☐ On dry basis

	<input checked="" type="checkbox"/> mol/mol fuel	<input type="checkbox"/> kg/kg fuel	Mole %	Mass %
CO <sub>2</sub>	1.080000	2.769057	6.2156	9.8222
CO				
H <sub>2</sub> O	2.446547	2.567756	14.0803	9.1082
N <sub>2</sub>	12.760381	20.825228	73.4381	73.8701
O <sub>2</sub>	1.088722	2.029672	6.2658	7.1995
SO <sub>2</sub>				
H <sub>2</sub>				
Ar				
unburned				
<b>TOTAL</b>	17.3757	28.1917	100.00 %	100.00 %

Dew Point (and P<sub>sat</sub>)

☐ 100%  mol/mol fuel  
 52.66 °C H<sub>2</sub>O (liquid)  %  kg/kg fuel  
 14.08 kPa



### Heat (Released or Absorbed)



**Exothermic**  
Exo = release

**ENERGY**

632.561 kJ/mol fuel

☐ Reactants

☐ Products

$$Q_f = \sum_{\text{Products}} n_{\text{out}} h_{\text{out}}(T_{\text{out}}) - \sum_{\text{Reactants}} n_{\text{in}} h_{\text{in}}(T_{\text{in}})$$

-796.836 kJ/mol fuel -164.275

$$Q_f = H_P(T_{\text{out}}, P) - H_R(T_{\text{in}}, P) = H_P(T_{\text{out}}, P) - H_P(T_{\text{ad}}, P)$$

☒ kJ/mol fuel ☐ kJ/kg fuel ☐ kW

$$H = H_{\text{form}} + H(T, P) - H(25^\circ\text{C}, 1 \text{ atm})$$

### Products Temperature

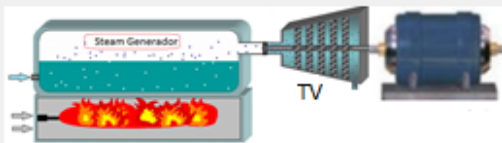
427.00 °C

### Adiabatic Flame Temperature

For  $\lambda = 1.6000$  1426.90 °C  
**(Actual)**

For  $\lambda = 1.0$  (max) 2017.90 °C (25°C, 1atm)  
**(Theoretical)**

### Boiler/Furnace Efficiency



74.21 %



☒ LHV

☐ HHV

### Stack heat losses



219.79 kJ/mol fuel

25.79 % (of the calorific value of the fuel)


### Results of the Energy Analysis





#### STEAM GENERATOR

Useful heat  kJ/mol fuel

Thermal efficiency  % 

Exhaust products loss  kJ/mol fuel

##### Mass ratio

Mol fuel/kg fluid

kg fuel/kg fluid

#### Thermal efficiency (Plant)

 % 

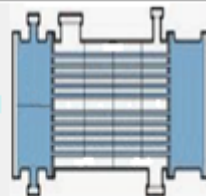
#### Net work

kJ/mol fuel

#### CONDENSER

*Released heat*

kJ/mol fuel



#### ADIABATIC PUMP

*Waste work*

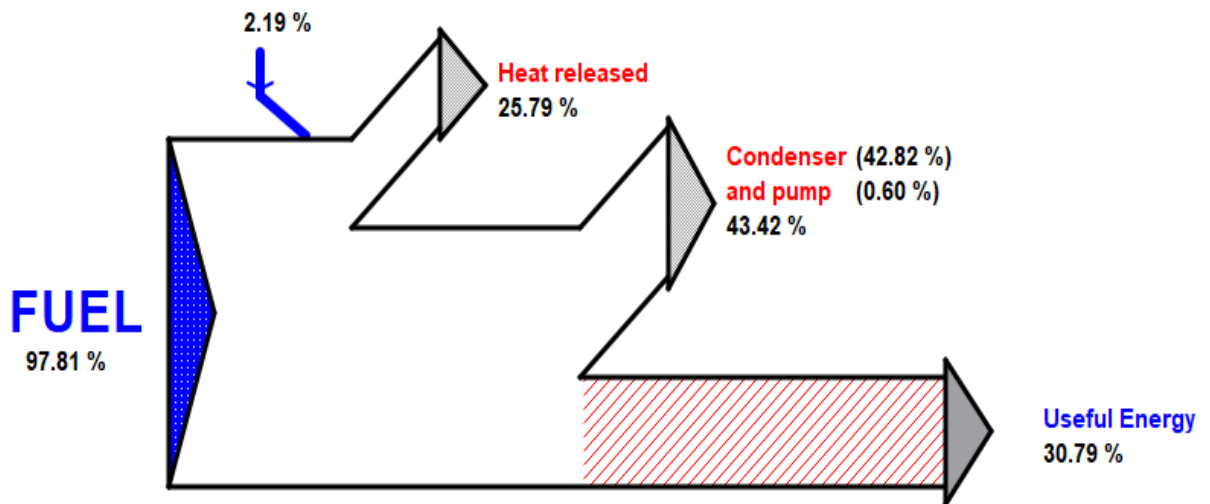
kJ/mol fuel

Isoentropic efficiency ?  %





ADIABATIC TURBINE			
	STAGE, HP	STAGE, MP (Alternative Exchanger)	STAGE, LP
Enthalpy at the entrance	549.09 kJ/mol fuel	560.77 kJ/mol fuel	573.10 kJ/mol fuel
<i>Specific work</i>	24.47 kJ/mol fuel	70.84 kJ/mol fuel	172.31 kJ/mol fuel
Stack discharge	524.62 kJ/mol fuel	489.93 kJ/mol fuel	400.79 kJ/mol fuel
Stack losses temperature	462.44 °C	315.36 °C	45.81 °C
Isoentropic efficiency	? 90.00 %	? 90.00 %	90.00 %
Isoentropic stack losses temperature	873.69 °C	550.00 °C	45.81 °C



### Method Exergy Applied to Chemically Reactive Systems



Exergy balances, which allow exergy destruction rates and exergy efficiencies to be determined, for the main components of the installation, that is, the boiler (differentiating between the combustion chamber and heat exchanger), turbine stages, condenser, and circulating pump.



Fuel chemical exergy

$b_{fuel}$   kJ/mol fuel

$b_{fuel}/LHV$

Chemical exergy (flue gas)

kJ/mol fuel

Exergy efficiency

**Steam generator**

?  %

**Useful exergy**  kJ/mol fuel

**Compressor**

?  %

**Turbine HP** **Turbine LP**

?  %  ?

**Turbine MP**

%

Sustainability index (SI)

$SI=1/D_p$   
 $D_p$ = total exergy destruction /exergy input

Reducing the environmental impact can be achieved by minimizing the irreversible exergy losses of the system.

Exergetic performance coefficient

(total exergy destruction/net work)





Irreversibilities

Combustion Chamber

**i** 312.15 kJ/mol fuel

Compressor

**i** kJ/mol fuel

Turbine (HP)

**i** 1.11 kJ/mol fuel

Heat Exchanger

**i** 141.27 kJ/mol fuel

Condenser

**i** 41.38 kJ/mol fuel

Turbine (MP)

3.38 kJ/mol fuel

Stack losses

**i** 108.80 kJ/mol fuel

Pump

2.23 kJ/mol fuel

Turbine (LP)

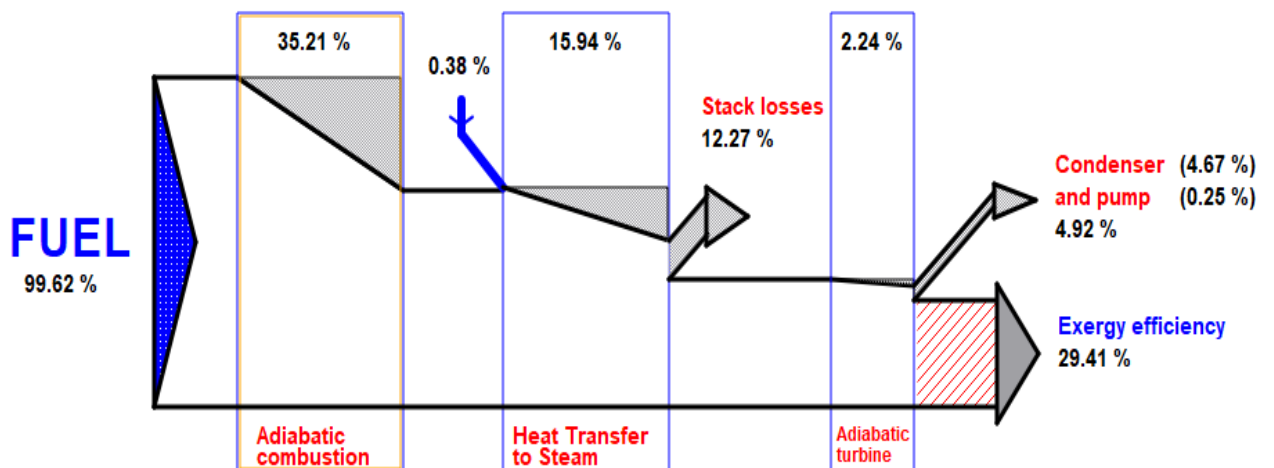
15.40 kJ/mol fuel

Regenerative Heat Exchanger

**i** kJ/mol fuel

**Exergy Destroyed (total)**

625.72 kJ/mol fuel



Si resolvemos el ciclo no regenerativo:


**Results of the Energy Analysis**





#### STEAM GENERATOR

Useful heat  kJ/mol fuel

Thermal efficiency  % 

Exhaust products loss  kJ/mol fuel

##### Mass ratio

Mol fuel/kg fluid

kg fuel/kg fluid

#### Thermal efficiency (Plant)

 % 

#### Net work

 kJ/mol fuel

#### ADIABATIC TURBINE

##### STAGE, HP

Enthalpy at the entrance  kJ/mol fuel

##### *Specific work*

 kJ/mol fuel

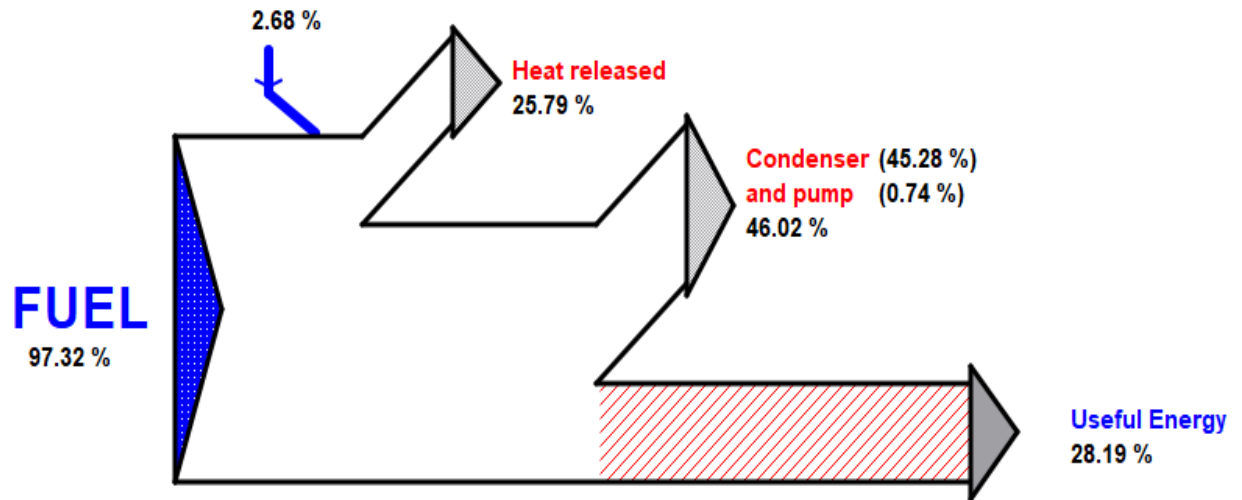
Stack discharge  kJ/mol fuel

Stack losses temperature  °C

Isoentropic efficiency ?  %

Isoentropic stack losses temperature  °C





### Method Exergy Applied to Chemically Reactive Systems

#### Irreversibilities

##### Combustion Chamber



312.15

kJ/mol fuel

##### Compressor



kJ/mol fuel

##### Turbine (HP)



25.62

kJ/mol fuel

##### Heat Exchanger



172.98

kJ/mol fuel

##### Condenser



25.58

kJ/mol fuel

##### Turbine (MP)



kJ/mol fuel

##### Stack losses



108.80

kJ/mol fuel

##### Pump



2.75

kJ/mol fuel

##### Turbine (LP)



kJ/mol fuel

##### Regenerative Heat Exchanger



kJ/mol fuel

#### Exergy Destroyed (total)

647.88

kJ/mol fuel

#### Sustainability index (SI)

1.37

SI = 1/Dp  
Dp = total exergy destruction  
/exergy input

Reducing the environmental impact can be achieved by minimizing the irreversible exergy losses of the system.

#### Exergetic performance coefficient

2.70

(total exergy destruction/net work)

